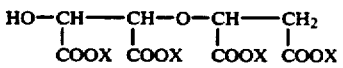
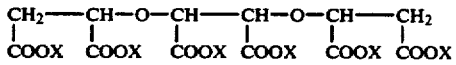
(*b*) *maintaining said aqueous reaction mixture at a temperature of from about 20° C. to 120° C. for a time period sufficient to form a reaction product mixture of*

(*i*) *tartrate monosuccinate of the formula:*



wherein X is a salt-forming cation; and

(*ii*) *tartrate disuccinate of the formula:*



wherein X is a salt-forming cation; and

(*c*) *reducing the calcium content of said reaction product mixture to the extent that the molar ratio of calcium to the tartrate succinate reaction products is less than about* 1:10.

22. A process according to claim 1 wherein the reaction mixture comprises from about 40% to 55% by weight of said salts of maleic and tartaric acids.

23. A process according to claim 22 wherein

(*a*) *the maleic acid salt is selected from sodium and potassium maleate;*

(*b*) *the tartaric acid salt is selected from sodium and potassium tartrate; and*

(*c*) *the calcium cations are provided by calcium hydroxide; and*

(*d*) *the monovalent cation-comprising neutralizing agent is selected from sodium hydroxide, potassium hydroxide and ammonium hydroxide.*

24. A process according to claim 23 wherein the maleic acid salt and tartaric acid salt reactants are formed in the reaction mixture in situ.

25. A process according to claim 23 wherein

(*a*) *the molar ratio of maleic acid to tartaric acid used in forming the reaction mixture of ranges from about 0.9:1 to 1.2:1 and*

(*b*) *the molar ratio of calcium cations to tartaric acid used in forming the reaction mixture ranges from about 0.8:1 to 1.5:1.*

26. A process according to claim 23 wherein the molar ratio of calcium to tartrate succinate reaction products is reduced to less than about 1:20 by addition the reaction mixture of a precipitating agent selected from alkali metal carbonate, alkali metal bicarbonate and mixtures thereof.

27. A process according to claim 23 wherein the aqueous reaction mixture is maintained at a temperature of from about 50° C. to 80° C. for a period of from about 0.5 to 10 hours.

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