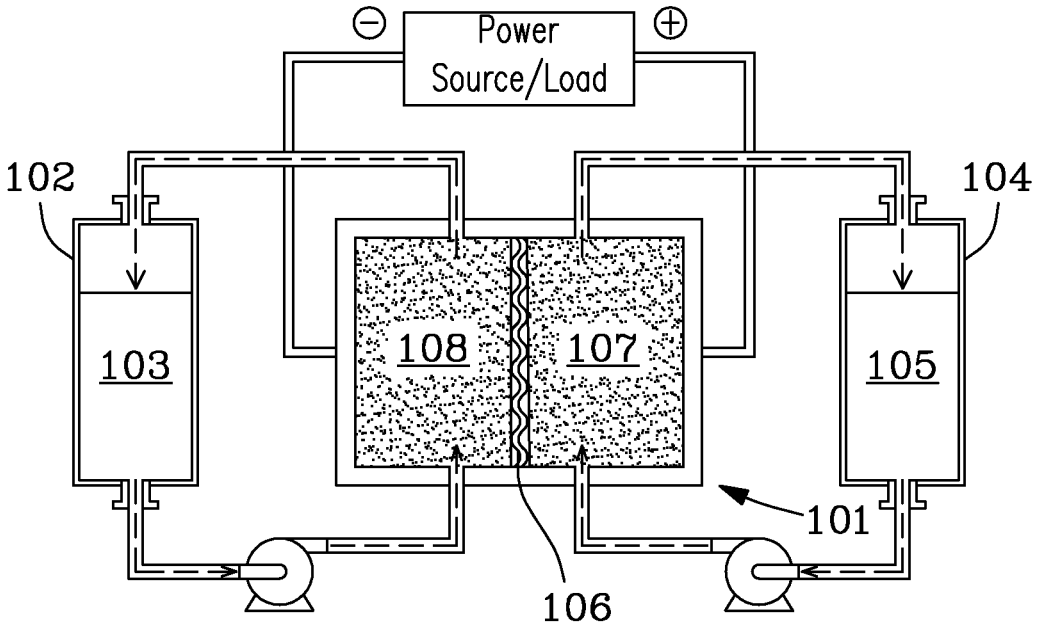

Fig. 1

Positive $Fe(CN)_6^{3-} + e^- \underset{\text{Charge}}{\overset{\text{Discharge}}{\rightleftharpoons}} Fe(CN)_6^{4-}$

Negative $\frac{1}{2}S^{2-} - e^- \underset{\text{Charge}}{\overset{\text{Discharge}}{\rightleftharpoons}} \frac{1}{2}S^*$

Overall $Fe(CN)_6^{3-} + \frac{1}{2}S^{2-} \underset{\text{Charge}}{\overset{\text{Discharge}}{\rightleftharpoons}} Fe(CN)_6^{4-} + \frac{1}{2}S^*$

* Element S can dissolve in excess Na_2S by forming sodium polysulfide solution (Na_2S_n)

Fig. 2